

**VISWAJYOTHI COLLEGE OF ENGINEERING AND TECHNOLOGY (Autonomous)**  
**FIRST SEMESTER B.TECH DEGREE EXAMINATION DECEMBER 2025 (2025 SCHEME)**

**Course Code: 25VJGXCYT122**

**Course Name: CHEMISTRY FOR INFORMATION SCIENCE / ELECTRICAL SCIENCE**

**Max. Marks: 60**

**Duration: 2 Hours 30 Minutes**

**PART A**

| <i>Q.No.</i> | <i>Answer all questions. Each question carries 3 marks</i>            | <i>CO</i> | <i>Marks</i> |
|--------------|-----------------------------------------------------------------------|-----------|--------------|
| 1            | Which are the advantages and disadvantages of SHE?                    | CO1       | (3)          |
| 2            | What is cathodic protection?                                          | CO1       | (3)          |
| 3            | Discuss the structure of Graphene.                                    | CO2       | (3)          |
| 4            | Write any three advantages of Super capacitors.                       | CO2       | (3)          |
| 5            | What is the principle of DETA?                                        | CO3       | (3)          |
| 6            | Give any three applications of IR spectroscopy.                       | CO3       | (3)          |
| 7            | What do you mean by sustainable development goals? List any two SDGs. | CO4       | (3)          |
| 8            | Explain disinfection by chlorination.                                 | CO4       | (3)          |

**PART B**

*Answer any one full question from each module. Each question carries 9 marks*

**Module 1**

- |    |                                                                                                                                                                                                                                                                                                                                                                    |     |     |
|----|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----|-----|
| 9  | a) Explain the conductivity measurement using Digital conductivitymeter.<br><br>b) Calculate the conductivity of given sample of water at 298K, which shows a conductance of 560 $\mu$ S in the given cell at 298K. A standard solution of 0.1 M KCl shows a conductance of 12.34 mS in that cell.<br>(Given that conductivity of 0.1M KCl at 298K is 0.01288S/cm) | CO1 | (9) |
|    | (6 marks)                                                                                                                                                                                                                                                                                                                                                          |     |     |
|    | (3marks)                                                                                                                                                                                                                                                                                                                                                           |     |     |
| 10 | Explain the mechanism of electrochemical corrosion of iron in acidic and alkaline environmental conditions .                                                                                                                                                                                                                                                       | CO1 | (9) |

**Module 2**

- |    |                                                                                                     |     |     |
|----|-----------------------------------------------------------------------------------------------------|-----|-----|
| 11 | What are nanomaterials? Explain the classification based on materials and dimension.                | CO2 | (9) |
| 12 | Explain the preparation of Polypyrrole. Write any three properties and applications of Polypyrrole. | CO2 | (9) |

**Module 3**

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|----|---------------------------------------------------------------------------------------------------------------------------------------------------|-----|-----|
| 13 | Explain the principle, instrumentation and applications of Scanning Electron Microscopy.                                                          | CO3 | (9) |
| 14 | a) Explain the instrumentation of UV-Visible Spectrophotometer. (6 marks)<br>b) Give any three applications of UV-Visible spectroscopy. (3 marks) | CO3 | (9) |

#### Module 4

- 15 a) What are Ion exchange resins? How exhausted resins are regenerated? ( 3 marks) CO4 (9)  
b) With a labelled diagram, explain the Ion exchange process used for demineralization of Hard Water. (6 marks)
- 16 Explain different stages in sewage water treatment with a flow diagram. CO4 (9)

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